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### Description

Electrolysis type method for producing water  
containing ozone, Apparatus thereof and Recovery  
method of solid polymer electrolyte membrane

## FIELD OF THE INVENTION

The present invention relates to a method for producing water containing ozone by water electrolysis, in detail, the present invention relates to the method and the same apparatus which can produce water containing ozone continually for long time. After here, the term of "water containing ozone" is shown by the term of "ozone aqua" .

## BACKGROUND ART

Conventionally, a gas dissolution method and water electrolysis method are known as typical methods for producing ozone aqua. Recently the water electrolysis method attracts attention, and practical use progresses. The gas dissolution method is a method producing ozone aqua by means of dissolving the ozone gas underwater. The water electrolysis method produces ozone aqua by means of the following process in which oxygen is generated at the side of an anodic electrode, the oxygen is converted to ozone by ozonization catalyst and the

ozone is immediately dissolved into water flowing in the anode side. In such a water electrolysis method, a small electrolytic apparatus is employed, which uses water available easily as a material and a small electric power supply having a specification of Several 10 Volt (V)  $\times$  several 10 amperes (A).

As ozone aqua production methods by a water electrolysis, there are inventions proposed by Japanese Patent Laid-Open No. 1-312092, No. 8-134677 and No. 8-134678. Summary of the apparatus will be described on the basis of a typical example shown in Fig. 17. In the Fig. 17, a solid polymer electrolyte membrane 5 disposes between an anode side casing 1 and a cathode side casing 2, which divides the anode side casing 1 and the cathode side casing 2 so as to be an anode chamber 6 and a cathode chamber 7 respectively. The solid polymer electrolyte membrane 5 is merely described with "membrane" or "electrolyte membrane" afterward. An inner surface of the anode side casing 1 is coating with materials having a corrosion resistance as against ozone such as a fluoro-resin or a glass. In the anode chamber 6 side, an anode 3 is pressed onto and touched with one face of the electrolyte membrane 5, which comprises noble metals 16 such as a platina having catalyst function for generating ozone. On the other hand,

in the cathode chamber 7 side, in similar, a cathode electrode 4 is pressed onto and touched with other face of the electrolyte membrane 5, which has a contact face consisting of noble metal 20 such as a platina or silver. Inflow openings 8,9 and outflow openings 10,11 for material water are formed by each of the anode chamber 6 and the cathode chamber 7. DC voltage is applied between both electrodes 3,4 from a direct-current power source 24 through electrode bars 19,23.

In the apparatus having such a constitution, while the water flows through the anode chamber 6 and the cathode chamber 7 respectively, the direct current is applied between both of the electrodes so as to energize. Then water electrolysis occurs across the electrolyte membrane 5. And oxygen and ozone generate in the anode 3 side, hydrogen generates in the cathode 4 side.

The ozone generated the anode 3 side becomes into ozone aqua to dissolve in water, then flows out from the outflow 10 as ozone aqua.

Here, wire nets 16,20 made with the noble metal such as a platina are employed to portions of the both electrodes 3,4 in contact with the electrolyte membrane 5 respectively in order to generate ozone aqua effectively as disclosed by a Japanese Patent

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Laid-Open No. 8-134677. In addition, on the back face portions of the both electrodes 3,4 in contact with the electrolyte membrane 5, lath nets 17,21 and electrode plates 18,22 are laminated sequentially and joined by mechanical conjugation methods such as a brazing, spot welding, other mechanical conjugation method so as to be integrated. The lath nets 17,21 are made of a materials such as titanium having the corrosion resistance as against ozone. The electrode bars 19,23 are joined with the electrodes having the integrated configuration as described above. While the material water flows through a duct which is formed by wire net and lath net, or lath net and lath net, intense turbulent flow and eddy current arise. By these flow and current, the ozone which arose in the anode side is instantly dissolved into water. Such a dissolution arises in the anode side continuously and then accumulates. As a result, the ozone aqua with high concentration is provided.

By the way, an ability of electrolyte membrane is getting down (membrane fouling) and a concentration of ozone aqua falls as time elapses while the apparatus as described above runs in succession in order to produce ozone aqua of predetermined concentration with an initialized

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current density. In order to prevent this problem, a method is taken, which raises current density. To describe in detail, as shown in Fig. 18, the current value  $A$  is controlled so as to maintain the concentration  $X$  of ozone aqua at a predetermined value  $X_s$  while keeping the current density uniformity. The current value rises as running time elapses, and finally, reaches a ceiling value  $A_e$  at a time  $t_1$ . The ceiling value  $A_e$  is the maximum value which apparatus can permit. In the condition that the current value has reached the ceiling value  $A_e$ , the current value cannot increase any longer. Therefore, the concentration  $X$  of ozone aqua decreases gradually, then falls in a predetermined threshold value  $X_e$  at the time  $t_2$ . When ozone concentration reaches this threshold value  $X_e$ , generally the running of apparatus is stopped. And the deteriorated electrolyte membrane 5 is exchanged after the apparatus is disassembled. However, according to this operational method, reassembling and disassembling of the ozone aqua production apparatus are troublesome and the apparatus operation efficiency becomes down remarkably, so that a cost for producing ozone aqua becomes expensive.

Therefore, one apparatus is proposed by this

applicant et al. in Japanese Patent Application No. 9-340188 (Japanese Patent Laid-Open No. 11-172482), wherein the electrolyte membrane 5 recovers easily without dismantling apparatus so as to reduce the complicated exchange work of the electrolyte membrane.

In this apparatus, the electrode bars 19,23 are respectively connected to in the back faces of the anode electrode 3 and the cathode electrode 4 having the lamination type structure as shown in Fig. 19. At this point, this apparatus is the same apparatus as shown in Fig. 17. However, the structure of the apparatus as shown in Fig. 19 is different from one as shown in Fig. 17 at the following point. In the apparatus shown in Fig. 19, the electrode bars 19,23 pass through through-holes 12,13 respectively. The through-holes 12,13 are respectively formed the anode side casing 1 and the cathode side casing 2. And edges of the electrode bars 19,23 are connected to fluid pressure cylinder devices 14,15 respectively. By this, the anodic electrode 3 and the cathode electrode 4 are able to advance and retreat as against the electrolyte membrane 5 respectively.

Operational method of this apparatus is described in the following.

In Fig. 18, when the current value A reaches to the

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allowed ceiling value  $A_e$  and the concentration  $X$  of ozone aqua reaches to the predetermined lower limit value  $X_e$ , the running apparatus is stopped (the supplies of the electricity and the water are cut off). And the fluid pressure cylinder 14,15 are actuated as shown in Fig. 20. Then both electrodes 3,4 are respectively apart from the electrolyte membrane 5, the pressing force against to the electrolyte membrane 5 is released. And, by means of keeping this condition for a regular time, the electrolyte membrane is recovered. The both electrodes 3,4 advance toward the electrolyte membrane again, and press the electrolyte membrane 5 by the predetermined pressing force. The supplies of the electricity and the running water are started, and the apparatus is started running again. In other words, the electrolyte membrane which has deteriorated as timewise during the apparatus running, is released from the pressing force and recovered. This operating state is shown in Fig. 22 with a time chart. In similar to the case in Fig. 18, the current value  $A$  is controlled so that the concentration  $X$  of ozone aqua may be maintained at the predetermined concentration  $X_s$  and rises as the running time elapses. At the time  $t_1$ , the current value  $A$  has reached to the ceiling value  $A_e$  allowed

by the apparatus. As the current value cannot rise any more, the supplies of the electricity and the running water to the apparatus are stopped and the running the apparatus stops. Both electrodes 3,4 are respectively kept apart from the electrolyte membrane 5 as shown in Fig. 20. when this condition is maintained for the predetermined time and the time reaches to the time  $t_3$ , the both electrodes 3,4 advance to again and press the electrolyte membrane 5. The supplies of the electricity and the running water start again. Then, the running apparatus is started again (at the time  $t_3$ ). Afterward, when the current value  $A$  reaches to the ceiling value  $A_e$  (at the time  $t_4$ ), the running apparatus is similarly stopped and, after the predetermined time elapses, the running apparatus is started again (at time  $t_5$ ).

These above operations are repeated. When the predetermined concentration of ozone aqua is not provided even if the stop and re-starting of the running apparatus are repeated, the electrolyte membrane 5 is exchanged.

In the above case, the number of times for exchanging the electrolyte membrane decrease and the life time of the electrolyte membrane improved. therefor the maintenance of the apparatus can be easy. In these points, this method has the



advantages in comparison with the prior method. However, as the apparatus must stop frequently, it is impossible to generate ozone aqua continuously. In this point, this method does not solve the conventional problems basically.

Therefore, the applicant proposes a method improving the above described method in the PCT application (PCT/JP98/5576) that insists on a priority based on Japanese Patent Application No. 9-340188. This method uses also the apparatus for producing ozone aqua by electrolysis as shown in Fig. 19 which has the mechanism for making the both electrodes advance or retreat against the electrolyte membrane. However, this method is different from the above-mentioned method.

To describe in detail, the concentration  $X$  of ozone aqua is kept at a constant value  $X_s$  as shown in a running time chart of Fig, 23, a process raising current value  $A$  is the same process as the above-mentioned method. When the current value  $A$  reaches to a ceiling value  $A_e$  which is a limit value allowed by the apparatus (time  $t_1$ ). A pressing force  $P$  of the electrodes for pressing the electrolyte membrane moves from an initial value  $P_1$  to the higher value  $P_2$  than  $P_1$ . Consequently, ozone generation rate improves and the current value  $A$  necessary to

maintain the predetermined ozone aqua concentration decreases. Accordingly, the current value  $A$  falls down from the ceiling value  $A_e$  to the normal operation value  $A_0$ . While the running is continued with this condition, the efficiency of the electrolyte membrane deteriorates again. The current value  $A$  necessary to maintain the predetermined concentration of ozone aqua increases gradually again. And the necessary current value reaches to the ceiling value  $A_e$  again (time  $t_6$ ). At this time point, the force value  $P$  for pressing to the electrolyte membrane moves to the higher value  $P_3$  than  $P_2$ . Consequently, the necessary current value  $A$  falls down from the ceiling value  $A_e$  to the normal operation value  $A_0$  again. Then, the current value  $A$  necessary to maintain the predetermined concentration of ozone aqua increases gradually again. And the necessary current value reaches to the ceiling value  $A_e$  again (time  $t_7$ ). At this time point, in case that the force value for pressing to the electrolyte membrane can increase more, the similar operations are repeated. However, in case that the force value for pressing to the electrolyte membrane reaches to the set ceiling value, the running apparatus is continued with the condition that the current value is the ceiling value  $A_e$ . And

, when the concentration  $X$  of ozone aqua decreases until under the lower limit value  $X_e$  (time  $t_8$ ), the running apparatus is stopped. And, the condition that the electrolyte membrane is pressed by the electrodes is released by a similar way to Fig. 20.

Then the stress which acts on the electrolyte membrane is released. Consequently, the membrane function is restored. After that, the following either process is done. One following process is that the electrolyte membrane is pressed by the electrodes again with the initial pressing force value  $P_1$ , and the supplies of the electricity current and the running water are also started again, then the running apparatus is started again. Other following process is that the apparatus is disassembled and the electrolyte membrane is exchanged, then the running of apparatus is started again.

According to this method, even if the necessary current value  $A$  reaches to the ceiling value  $A_e$ , the necessary current value can falls down by means of making the force  $P$  for pressing to the electrolyte membrane by the electrodes increases. Consequently, the running continuation clock time of the apparatus gets longer by leaps and bounds. However, when the pressing force  $P$  by the electrodes reaches to the



the side of anode by means of water electrolysis to supply water onto both sides of the electrolyte membrane, further, a pressing force condition varies with a preset condition in the state that water and electricity are continuously supplied, and then, returns to the initial pressing force condition, the pressing force condition means one that the anodic electrode, the cathode electrode or the both electrodes press against the electrolyte membrane. Thereby, the electrolyte membrane recovers while the ozone aqua is produced.

In this method, the operation for changing the pressing force against the electrolyte membrane can be carried out by means of decreasing or increasing the pressing force from the electrodes or the combination of these.

In addition, a following pattern for changing the pressing force is preferable, which includes keeping a state lower or upper than the preset pressing force for a regular time. Furthermore, this operation can be done by following two ways. In one way, this operation repeats with the predetermined cycle. In other way, this operation is done once or plural times when it is detected that an ability of producing ozone aqua falls down under the predetermined value. It is optional which

[illegible][illegible][illegible]

when it is detected that the concentration of ozone aqua or the producing ability deteriorates under the predetermined value, such an operation is.

Besides, in the first and second methods, the electrolyte membrane will not be recovered enough even if the operation for varying the pressing force, or the operation for varying the current or voltage forcibly is carried out. In that case, the production of ozone aqua is stopped and the both electrodes are respectively apart from the electrolyte membrane. This state is kept for the predetermined time then the electrolyte membrane recovers. After that, the production of ozone aqua is started again with the predetermined condition about the operating. According to this method, the life time of the electrolyte membrane can be improved all the more.

Furthermore, there is another way. In this method, when the electrolyte membrane will not be recovered enough even if the operation for varying the pressing force or the operation for varying the current or the voltage forcibly is done, the pressing force value of the electrode against the electrolyte membrane moves to the high pressing force value which is set previously. And the production of ozone aqua is continued. According to this method, it is able to produce ozone aqua continuously for a long time

by leaps and bounds.

#### BRIEF DESCRIPTION OF DRAWINGS:

Fig. 1 shows a running time chart in an embodiment concerning the first ozone aqua production method of the present invention.

Fig. 2 shows a flow chart concerning an operation control method of Fig. 1.

Fig. 3 shows a running time chart in a modification example of a variation of pressing force concerning the method of Fig. 1.

Fig. 4 shows a running time chart in other embodiment concerning the first method of the present invention.

Fig. 5 shows a running time chart in a furthermore other embodiment concerning the first method of the present invention.

Fig. 6 shows a flow chart in the operation control method of Fig. 5.

Fig. 7 shows a running time chart in a furthermore other embodiment concerning the first method of the present invention.

Fig. 8 shows a running time chart in a furthermore other embodiment concerning the first method of the present invention.

Fig. 9 shows a flow chart in an operation



control method of Fig. 8.

Fig. 10 shows a running time chart in a furthermore other embodiment concerning the first method of the present invention.

Fig. 11 shows a running time chart in an embodiment concerning the second method of the present invention.

Fig. 12 shows a feature section of an electrolysis type ozone aqua production apparatus in an embodiment of the present invention.

Fig. 13 shows a feature section of an electrolysis type ozone aqua production apparatus in another embodiment of the present invention.

Fig. 14 shows a feature section of an electrolysis type ozone aqua production apparatus in furthermore another embodiment of the present invention.

Fig. 15 shows an example of a running time chart in a real running concerning the first method of the present invention.

Fig. 16 shows an example of a real running time chart in a conventional method.

Fig. 17 shows a feature section of an example of a conventional electrolysis type ozone aqua production apparatus.

Fig. 18 shows a running time chart in a running

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example of the apparatus in Fig. 17.

Fig. 19 shows a feature section of a conventional electrolysis type ozone aqua production apparatus and a feature section of another example of an electrolysis type ozone aqua production apparatus using in the present invention.

Fig. 20 shows a feature section of recovery state of a solid polymer electrolyte membrane according to a conventional apparatus of Fig. 19.

Fig. 21 shows a conception diagram of a control system in an electrolysis type ozone aqua production apparatus of the present invention.

Fig. 22 shows a time chart of a conventional operational method in the apparatus of Fig. 19.

Fig. 23 shows a time chart of a conventional other operational method in the apparatus of Fig. 19.

Fig. 24 shows feature section of other example of an electrolysis type ozone aqua production apparatus of the present invention.

Fig. 25 shows a feature section of furthermore other example of an electrolysis type ozone aqua production apparatus of the present invention.

The best mode for carrying out the invention

The present invention will be described in detail by drawings as follows. At first, the first

method of the present invention will be described. As shown in Fig. 21, an ozone aqua production apparatus in this method comprises a main body 1 of the ozone aqua production apparatus, a pressing force controller 81, a setting adjuster 82 of pressing force, an electric power supply 24, a material water feeder 83, a concentration detection sensor 84 of ozone aqua having a detector and a main controller 85 of the ozone aqua production apparatus. Here, in the main body 1 of the ozone aqua production apparatus, the portion having the same symbol as in the apparatus of Fig. 19 shows the same constitution as in Fig. 19. So, a repeating explanation is omitted.

The pressing force controller 81 is connected to an advance or retreat drive parts 14,15 of the main body 1, and controls pressing forces from an anodic electrode 3 and a cathode electrode 4 toward an electrolyte membrane 5. The setting adjuster 82 of pressing force is connected to the pressing force controller 81, and able to set in a predetermined pressing force. The electric power supply 24 is a power source to apply voltage to the both electrodes 3,4, and comprises a current detector 91 for detecting a current value. The material water feeder 83 supplies material water for generating



detection sensor 84 with a desired value ( $X_s$ ) stored in the storage part 89 at the time of transmitting.

And the first comparator 88a outputs a signal showing a deviation of the ozone aqua concentration toward the first command section 86a. On the other hand, when the measured value ( $X$ ) reaches the desired value ( $X_s$ ), the first comparator 88a outputs an attainment signal to the second command section 86b.

The second comparator 88b is connected to the current detector 91 of the electric power supply 24, the storage part 89 and the third command section 86c. The second comparator 88b compares the ceiling value ( $A_e$ ) of the electric power supply, which is stored by the storage part 89, with a detected current value ( $A$ ) detected by the current detector 91 of the electric power supply 24, then, when the detected current value ( $A$ ) reaches the ceiling value ( $A_e$ ), outputs an attainment signal toward the third command section 86c.

The third comparator 88c is connected to the timer 87, the storage part 89, the first command section 86a and the third command section 86c. The third comparator 88c compares an elapsed time ( $t$ ), which is an elapsed time measured by timer 87 during running or stopping the apparatus, with such a predetermined time ( $T_c$ ) as preset and stored in the

storage part 90, then, when the running time of apparatus or the stopping time (t) reaches the predetermined time ( $T_c$ ), outputs an attainment signal toward the first command section 86a or the third command section 86c.

And, the fourth comparator 88d compares a stop frequency (N), which is counted by the stop frequency count part 90 for measuring the frequency that apparatus stopped, with such a predetermined stop frequency ( $N_e$ ) as preset and stored in the storage part 89, then outputs the resultant toward the third command section 88c.

In the storage part 89, the various necessary preset values for running is memorized and stored, which includes a production target concentration ( $X_s$ ) of ozone aqua which is previously set as an operating condition, a control start concentration ( $X_m$ ) for starting a predetermined control as described below, a permission lower limit concentration ( $X_e$ ), a control cycle time ( $T_c$ ) of the apparatus as described below, a variation pressing force retention time ( $T_l$ ) as described below and a ceiling value of a current value ( $A_e$ ) etc.

The first command section 86a of the control part 86 outputs an order toward the electric power supply 24. The order is an order which varies a

current value (A) supplied between the both electrodes 3,4 according to a signal from the first comparator 88a comparing an ozone aqua concentration so that the ozone aqua concentration (X) may approach to a desired value (Xs).

The second command section 86b is connected to the first comparator 88a, the third comparator 88c and the pressing force controller 81. The second command section 86b outputs an order for varying the pressing forces of the both electrodes 3,4 toward the pressing force controller 81 when it is detected that an elapsed time in state of running the apparatus (t) reaches the setting value (Tc) according to the signal from the third comparator 88c. And, the second command section 86b outputs an order for varying pressing forces of both electrodes 3,4 toward the pressing force controller 81 when it is detected that the concentration (X) of ozone aqua reaches the preset value of the control start concentration (Xm) according to the signal from the first comparator 88a. Further, the second command section 86b outputs an order which exchanges the pressing force of both electrodes 3,4 with an initial pressing force, toward the pressing force controller 81 when it detects that the elapsed time (t) reaches the above-mentioned predetermined value (T1)

according to the signal from the third comparator 88c.

The third command section 86c is connected to the first comparator 88a, the second comparator 88b, the third comparator 88c, the electric power supply 24, the material water feeder 83, the pressing force controller 81 and the timer 87. The third command section 86c outputs an order for turning off the electric power supply 24, the material water feeder 83 and a pressure regulator 21 when it detects that a current value reaches the ceiling value (Ae) according to the signal of the second comparator 88b and when it detects that the concentration of ozone aqua reaches a lower limit value (Xe) according to the signal from the first comparator 88a. Further, the third command section 86c outputs an order which works the apparatus again when a predetermined time passes according to the signal from the timer 87, toward the electric power supply 24, the material water feeder 83 and the pressing force controller 81.

Next, an embodiment on the first method of the present invention will be explained using such the drawings as Fig. 1 showing a running time chart, Fig. 2 and Fig. 21 showing a flow chart for executing such the run as shown in the time chart of Fig. 1.

At first, in Fig. 2, a main switch of the



electric power supply 24 is turned ON at start-up. Then, a supply of electricity to a control system begins, a supply of material water begins at inflow openings 8,9. DC voltage is applied between the both electrodes 3,4 (S 1). Next, the pressing forces of both electrodes 3,4 toward the electrolyte membrane 5 is set to an initial value P1 by the setting adjuster 82 (S2). And, a measurement of an elapsed time t is started by the timer 87 (S 3). When ozone aqua is produced by electrolysis and starts flowing out from the outflow opening 10, the concentration detection sensor 84 starts and continues always detecting a concentration X of the ozone aqua. The first comparator 88a compares the detected concentration X with the desired concentration value Xs of ozone aqua (S 4). Here, the desired value Xs is set so as to permit a predetermined dispersion (x), therefore the detected concentration X is really compared with  $Xs \pm x (\doteq Xs)$ . When the concentration of ozone aqua is lower ( $X < Xs$ ) (S 4, No), the first command section 86a puts out a command toward the electric power supply 24. The command is a command for putting up from A to A+a to add a preset current value of a to the recent current value A. And the operation is done with a new current value (S 5). On the other hand, when the

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concentration  $X$  of ozone aqua is higher than the predetermined desired value ( $X > X_s$ ) (S 4, No), the first command section 86a puts out an instruction signal toward the electric power supply 24. The instruction signal is an instruction signal for changing the current value to  $A-a$  to reduce the current value from the recent current value  $A$  by  $a$ . In this manner, the current value  $A$  is controlled so that the concentration  $X$  of ozone aqua may be always maintained at the vicinity value of the  $X_s$ .

Fig. 1 shows an advanced state after the concentration of ozone aqua reached  $X_s$  in this manner. During continuing the operations with this state, the values  $A, V$  of the current and voltage rise gradually as the time  $T$  elapses as shown in Fig. 1 so as to maintain the concentration  $X$  of ozone aqua at the concentration  $X_s$  as against a deterioration of the electrolyte membrane. Therefore, a suitable time ( $T_c$ ) is previously set so that the concentration  $X$  of ozone aqua may not fall under the allowed lower limit value  $X_e$  because of a reduction of an efficiency of the electrolyte membrane. The third comparator 88c watches that the running time  $t$  elapses. Then, when the elapsed running time  $t$  reaches a state of  $t \geq T_c$  (time  $t10$ ) (S6, Yes), the pressing force controller 81 is controlled with

keeping the supply of the electricity and the water (in other words, with continuing the operation of the apparatus) according to a signal of the second command section 86b so that the pressing forces P of the both electrodes 3,4 may vary from the initial value P1 to the low value P4 (S7). And, while the pressing force P changes, the timer begins measuring the elapsed time (t') after starting changing of the pressing force (S8).

For the reason that the contact resistance between the electrolyte membrane 5 and both electrodes 3,4 increases by the reduction of pressing force, the current value A falls down to a current value A1 lower than the initial value A0 and also the concentration X of ozone aqua begins falling from the specified value Xs slightly. On the other hand, the voltage V rises to the maximum voltage value Ve of the power source. In this state, the concentration of ozone aqua continues falling, however, this state is held during a predetermined period of time (T1) until a suitable time (time t11) which falls in the period of time in which the concentration of ozone aqua is maintained in the same or upper value than the permission lower limit value Xe. In other words, the third comparator 88c outputs a signal to the second command section 86b (S9, Yes)



3,4 against to the electrolyte membrane are maintained at the low pressing force  $P_4$ . The recovery principle is not clear. However, it is inferred as follows. Deleterious materials ( For example, Ion of impurities contained in hydrogen ion and the material water ) accumulate in the membrane or on the membrane face as the time elapses as described below.

By means of giving the change of the pressing force against the electrolyte membrane of which function falls down by the deleterious material, an equilibrium state of the deleterious materials is damaged. Consequently, the recovery concerning the ability of the membrane is accomplished.

Similarly, when the elapsed time reaches to a cycle  $T_c$  (time  $t_{12}$ ), by the above-mentioned knack, the pressing force  $P$  falls from the initial value  $P_1$  to the value  $P_4$ . After keeping this state for the fixed period  $T_1$  until a time  $t_{13}$ , the pressing force returns to the initial value  $P_1$  again. Such operations are repeatedly performed without stopping the apparatus,

In the above method, the period  $T_1$  in which the pressing force is held at the low value  $P_4$  (hereinafter, it is called as "the low period" ) is basically voluntarily determined until the

concentration of ozone aqua reaches the allowed lower limit value  $X_e$ . Also, the repetition cycle ( $T_c$ : the period from  $t_{10}$  to  $t_{12}$ ) is determined until the concentration of ozone aqua reaches the allowed lower limit value  $X_e$ . However, a relation of both must be considered. In other words, when the repetition cycle  $T_c$  is long, the low period  $T_l$  becomes long. And when the repetition cycle  $T_c$  is short, the low time  $T_l$  shortens. Generally, when the repetition cycle  $T_c$  is set in a range of 10 - 30 minutes, the low period  $T_l$  is set in a range of several seconds, i.e. 1-5 seconds.

In addition, when the recovery ability of the electrolyte membrane by means of varying the pressing force nears to the limit, the current value  $A$  will become the high value than an initial value  $A_0$  even if the pressing force returns the initial value  $P_1$ . Accordingly, the current value  $A$  becomes high gradually at every operation for changing the pressing force. At last, the electric current reaches the ceiling value  $A_e$  for only a short time and the concentration  $X$  of ozone aqua reaches under the permission lower limit value  $X_e$ . Advanced treatments which should be taken after such a state occurs, will be described below.

In Fig. 1 as mentioned above of, the pressing

force  $P$  suddenly changes from the initial value  $P_1$  to the low value  $P_4$  then the low state is maintained for the regular period ( $T_1$ ). As to the change pattern of the pressing force, there is various kinds of modification examples. Fig. 3 shows the various kinds of examples as to the change pattern of the pressing force.

In Fig. 3(a), the pressing force  $P$  varies from the initial value  $P_1$  to a value  $P_5$  lower than  $P_4$  with V-shaped pattern. In this case, the pressing force is maintained at the value lower than  $P_4$  during the predetermined period of times  $t_{16}$ - $t_{17}$ . And, as the pressing force  $P$  falls down, also the electric current  $A$  falls down to a value  $A_2$  lower than the low electric current  $A_1$  of the case of Fig. 1. Then, as the pressing force  $P$  increases, the current value  $A$  rises gradually. On the other hand, as the pressing force  $P$  falls down, the voltage  $V$  rises up. When the voltage  $V$  reaches the permission maximum value  $V_e$ , the state is maintained. Then, as the pressing force recovers, the voltage  $V$  falls down to the initial value  $V_0$ .

Next, Fig. 3(b) differs from Fig. 3(a) in the point that the pressing force  $P$  varies with U-shaped pattern. The pressing force is kept at a value lower than the set low value  $P_4$  during the period of

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t18-t19. Further next, Fig. 3(c) is same as Fig. 3(a) in the point that the pressing force changes with V-shaped pattern. However, as soon as the pressing force reaches the predetermined low pressure  $P_4$  at the time  $t_{20}$ , the pressing force rises to the value  $P_1$  immediately. In other words, in this case, there is no the period in which the pressing force is maintained with the predetermined value lower than the low value  $P_4$ . But, there is no noticeable problem with this method either. However, in this case, there is a tendency in which the recovery ability of the electrolyte membrane is inferior to such the operations as the operation for changing the pressing force as shown in Fig. 1 and the operation for changing the pressing force of Fig. 3(a) and Fig. 3(b). Accordingly, a consideration is necessary, such as a shortening the repetition cycle or setting the low pressure value  $P_4$  further low.

Next, Fig. 4 shows other embodiment with regard to the method of Fig. 1. The pressing force  $P$  falls down from the initial value  $P_1$  to the low value  $P_4$  at a suitable time  $t_{21}$  which falls in the period where the reduction of the efficiency of the electrolyte membrane is a little during running the apparatus. The above operation is same in case of Fig. 1. In the present embodiment, the return and





the third comparator 88c and compared with each other. The resultant of the comparison is output to the third command section 86c at every  $\Delta t$ . And, the change order of the pressing force P is output to the pressing force controller 81 from the third command section 86c.

In this case, with regard to the change of the pressing force, the repetition period (T 2) and the operation cycle (Tc: the period of  $t_{21}$ - $t_{23}$ ), similarly to the above said case, are optionally determined in the period by the concentration X of ozone aqua falls down to the permission lower limit value  $X_e$ . In addition, when the efficiency of the electrolyte membrane will not recover enough even if the operation for changing the pressing force P is carried out, the concentration X of ozone aqua will reach the permission lower limit value  $X_e$  finally. An advanced treatment after this state will be described later.

Next, Fig. 5 shows a running time chart as to another embodiment of the present invention. While the running is continued to be controlled so that the concentration of ozone aqua may reaches the predetermined desired value  $X_s$ , the current value A rises gradually and reaches the ceiling value  $A_e$  allowed by the apparatus at the time 31. When the

running is further continued with this state, the concentration of ozone aqua falls down gradually. Therefore, a suitable value  $X_m$  is previously set between the permission lower limit value  $X_e$  and the specified value  $X_s$ . When this preset value  $X_m$  is detected (time  $t_{32}$ ), as similar to the case of Fig. 1, the pressing force of the electrolyte membrane varies from the initial value  $P_1$  to the low pressure value  $P_4$ . Then, this state is kept during the predetermined period until time  $t_{33}$  ( $T_3$ ). After that, the pressing force returns to the initial value  $P_1$  again. In other words, in the examples of Fig. 1, Fig. 3, Fig. 4, the recovery of the electrolyte membrane is accomplished at every predetermined cycle ( $T_c$ ) by means of carrying out the operation for changing the pressing force. On the other hand, in the present embodiment, the operation for changing the pressing force  $P$  is begun to detect the decrease of the concentration of ozone aqua. In this point, the present embodiment differs from the above mentioned embodiments. But, the recovery of the electrolyte membrane is accomplished by the operation for changing the pressing force  $P$ . In this point, the above mentioned embodiments are same as the present embodiment.

Fig. 6 shows a flow chart which realizes this

running time chart.

In Fig. 6, the concentration of ozone aqua is controlled so as to become the predetermined value  $X_s$  till the current value  $A$  reaches the ceiling value  $A_e$  of the apparatus. As concerns these points (S1-S4), the present embodiment is the same case of Fig. 2. When the second comparator 88b detects that the current value  $A$  has reached the ceiling value  $A_e$  (S5, Yes), the concentration of ozone aqua begins decreasing. Accordingly, the concentration detection sensor 84 of ozone aqua detects continuously the decrease condition as concerns the concentration of ozone aqua, When the concentration  $X$  of ozone aqua decreases toward and under the preset value  $X_m$  for starting controlling the change of the pressing force (S7, Yes), the second command section 88b puts out the instruction for changing the pressing force  $P$  to the low value  $P_4$  (S8). The third comparator determines whether the elapsed time  $t'$  has elapsed by the predetermined period  $T_3$  after the pressing force begins changing (S8, S9). When this elapsed time  $t'$  fulfills  $t' \geq T_3$  (S10, Yes), the third comparator outputs the exceeding the signal to the second command section 86a. And, the second command section 86b outputs the signal for setting the pressing force  $P$  at the initial value  $P_1$  again toward

the pressing force controller 81 ( Shifting to S2).  
Then, the running apparatus is continued.

In this method, the V-shaped or U-shaped pattern as shown in Fig. 3 or other pattern can be taken as the pattern for changing the pressing force. Also, in the operation period (T3), it is available that the operation for changing the pressing force with the pulse-shaped pattern is carried out at every minute time  $\Delta t$  and plural times as similarly to the method shown in Fig. 4. Further, in the present embodiment, when it is detected that the concentration of ozone aqua has reached  $X_m$ , the operation for changing the pressing force is started.

However, the operation for changing the pressing force may be started at the event when it is detected that the electric current or voltage or a quantity of supply of the material water has reached the predetermined value (There is such a method as the concentration of ozone aqua is recovered by decreasing a quantity of supply of material water when the concentration of ozone aqua decreases.).

Next, Fig. 7 shows a running time chart with regard to another embodiment of the present invention. There is the different point from the

method of Fig. 1 - Fig. 5 in the operation for changing the pressing force  $P$  against the electrolyte membrane. In other words, in the method of Fig. 1, and Fig. 3-5, the change of the pressing force  $P$  is decrease. On the other hand, in the present embodiment, the change of the pressing force  $P$  is increase. In detail, in Fig. 7, when the running is performed so that the concentration of ozone water may be maintained at the predetermined value  $X_s$ , the current value  $A$  increases gradually as previously described. The pressing force of the electrodes against the electrolyte membrane increases from the initial value  $P_1$  to the higher value  $P_6$  at the suitable time ( $t_{41}$ ) before the current value  $A$  reaches the permission ceiling value  $A_e$  of the apparatus. And, the electrolyte membrane is recovered by holding this state during the regular period  $T_4$  until the time  $t_{42}$ . In this case, as the contact resistance between the both electrodes and the electrolyte membrane decreases, the current value  $A$  rises till the ceiling value  $A_e$ , the voltage falls down to the low value  $V_1$  conversely and the concentration of ozone water slightly rises for a short time. However, when the pressing force returns to the initial value  $P_1$  after the regular period ( $T_4$ ) elapsing, the current value  $A$ , the

Case	Age	Sex	Occupation	Duration of Illness	Onset	Course	Outcome
1	45	M	Farmer	10 years	1955	Chronic	Death
2	52	F	Homemaker	5 years	1960	Chronic	Death
3	68	M	Retired	3 years	1965	Chronic	Death
4	72	F	Teacher	2 years	1968	Chronic	Death
5	78	M	Engineer	1 year	1970	Chronic	Death
6	82	F	Homemaker	6 months	1972	Chronic	Death
7	85	M	Retired	4 months	1973	Chronic	Death
8	88	F	Homemaker	3 months	1974	Chronic	Death
9	90	M	Retired	2 months	1975	Chronic	Death
10	92	F	Homemaker	1 month	1976	Chronic	Death

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7	85	M	Retired	4 months	1973	Chronic	Death
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7	85	M	Retired	4 months	1973	Chronic	Death
8	88	F	Homemaker	3 months	1974	Chronic	Death
9	90	M	Retired	2 months	1975	Chronic	Death
10	92	F	Homemaker	1 month	1976	Chronic	Death

running continuous for long term. Fig. 8 shows a running time chart in case that the operation for changing the pressing force is repeated and the running apparatus continues by the method as shown in Fig. 1. In Fig. 8, an interval (a) shows a state in normal running, wherein the pressing force decreases from the initial value  $P_1$  to the low value  $P_4$  with the definite cycle according to mentioning above, and the state is held during the regular period ( $T_1$ ), then the pressing force returns to the initial value  $P_1$ . This operation is repeated. The concentration  $X$  of ozone aqua is always maintained at the desired value  $X_s$ .

However, when this running is continued for the long term, the recovery ability of the efficiency of the electrolyte membrane deteriorates gradually, as indicated in an interval (b) of Fig. 8, the current value  $A$  reaches the permission ceiling value  $A_e$  of the apparatus before the elapsed time reaches the time  $t_{64}$  at which the operate for decreasing the pressing force begins.

When this state continues further, as indicated in an interval (C) of Fig. 8, the current value reaches the ceiling value  $A_e$  at time  $t_{65}$ , also the concentration of ozone aqua begins deteriorating from this time point then the



recovering the function of the electrolyte membrane will not be enough even if the operation for changing the pressing force P is carried out.

Therefore, the current value A does not decrease from the ceiling value  $A_e$  and the tendency of the ozone aqua concentration to decrease continues. Finally, the concentration of ozone aqua reaches the lower limit value  $X_e$  at the time  $t_{66}$ . In such a state, it is meaningless to continue the operation for recovering the electrolyte membrane by changing the pressing force P.

At this time, in the present invention, there are three ways to be taken. The first way is a way wherein the electrolyte membrane is exchanged after the apparatus is stopped and disassembled. The second way is a way shown in Japanese Patent Laid-Open No. 11-172482 wherein it is waited that the membrane function recovers after the apparatus is stopped and the working of the electrolyte membrane is stopped. The third way is shown in PCT/JP98/5576 wherein the pressing force to the electrolyte membrane increases and a membrane is revitalized. Hereon, the second and the third way will be described. A detailed description about the first way is omitted because the electrolyte membrane is only exchanged in the first way.



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P1 again and the running is continued (Shift to S2).

Next, when the state reaches the state (c) of Fig. 8(c) (S5), the concentration X of ozone aqua decreases gradually after the current value A reaches the ceiling value Ae. And, when the concentration X of ozone aqua is less than the lower limit value Xe (S 7, No), the signal from the first comparator 86a is sent to the third command section 86c. Here, the power source is set to OFF, also the supply of material water stops (an interval d of Fig. 8), and the anodic electrode 3 and the cathode electrode 4 are kept apart from the electrolyte membrane 5 as indicated in Fig. 20, then the run of the apparatus stops (S12). This stop process is equivalent to the membrane recovery process described in an above-mentioned Japanese Patent Laid-Open No. 11-172482, wherein, the work of the electrolyte membrane 5 is stopped and the recovery as to the function of electrolyte membrane 5 is carried out. The stop frequency N about the run is counted by the stop frequency count part 90 (S13). It is judged whether the stop frequency reaches the predetermined frequency Ne in the fourth comparator 88d. When the stop frequency does not reach the predetermined frequency Ne (S14, No), the judgment is continued (S16, No) at the third comparator 88c

until the stopping time  $t$  elapses by the predetermined stop period  $Tr$ . When the predetermined stop period  $Tr$  elapses (S16, Yes), the elapsed signal is be transmitted to the third command section 86c then, the run of the apparatus is started (an interval  $e$  of Fig. 8, Shift to S1).

In the stop period  $Tr$  of the above apparatus, the pressing force to the electrolyte membrane is completely released. In this interval, the impurities etc. which have accumulated in the electrolyte membrane by the pressing force, are given off, then the membrane recovers. As this stop period  $Tr$ , the time is necessary at least more than 30 minutes. It is preferable about 3-12 hours. When the function of the electrolyte membrane has recovered in this stop period, the electrolyte membrane is pressed again by the both electrodes from the time  $t67$  with the initial value  $P1$  as shown in the interval (e) of Fig. 8. And the supplies of the material water and the electricity are begun again.

When the running is carried out again with the operating condition in conformity to the initial setting, the current value  $A$  starts from the initial value  $A0$  in similar to the case of Fig. 1. The concentration  $X$  of ozone aqua reaches the predetermined value  $Xs$  immediately. Afterward, as

shown in Fig. 1, the operation for changing the pressing force is repeated and the running is continued. And, when the efficiency of the electrolyte membrane deteriorates again and the predetermined concentration of ozone aqua is not provided, the recovery operation of the membrane function is carried out by stopping the run again and releasing the pressing force to the electrolyte membrane according to the statement above.

Afterward, such the operation is repeated as well, finally, the recovery of the membrane function cannot be expected. Therefore, It is judged whether the stop frequency N reaches the predetermined frequency Ne in the fourth comparator 88d. The stop frequency N is counted in the stop frequency count part 90 until then (In other words, the recovery frequency of the electrolyte membrane by stopping the run of the apparatus). When the stop frequency N reaches the predetermined frequency Ne (S14, Yes), the electrolyte membrane is exchanged as prior (S17). And, the stop frequency N is reset to zero ( $N = 0$ ) (S 18).

Next, the third above way will be explained by using Fig. 10. The way of Fig. 10 is also one example of the way in which the operation for changing the pressing force of Fig. 1 is carried out.



[illegible][illegible][illegible]

[illegible][illegible][illegible]



larger than one of positive charge then effective dipole layer is formed on the ion clusters surfaces.

It is the reason why the distribution width of negative charge becomes larger than one of positive charge that asymmetry forces act between the negative fixed charges by the arrangement energy of polymer chain. When ions transfer among clusters by jumping, the ions must be over potential barrier which the dipole layers make. This potential barrier becomes large as size or number of ion clusters becomes large, and the number and size of ion clusters in the electrolyte membrane become large as time elapses. Therefore, it is inferred that transmission barrier against ions become large as the time elapses. Especially, a lot of impurities are contained in well water or city water used for producing ozone aqua, so, when they are electrolyzed, the ions having various potentials transmit through the membrane. And it is inferred that, the ions easy to transmit and the ions hard to transmit, simultaneously act on the electrolyte membrane, so that, the ion clusters are accumulated in the membrane, then, as the whole, the transmittable area becomes small gradually as the time elapses. And it is inferred that, in this state, a physical shock is given to the membrane by changing

the pressing force then equilibrium state of ion clusters are damaged, so that, the membrane function revives.

On the other hand, the second hypothesis is given as follows. The change of the pressing force is given to the electrolyte membrane in which the transmission resistance of the ions are increased by the polarization effect. Thereby, cancellations of the polarization states are carried out. In other words, in this hypothesis, when the water electrolysis is carried out, hydrogen ions transmit through the cathode side. But, because of the a poor solubility, the hydrogen ions become hydrogen moleculars by ionic bond then fall in gas phase. As the part, through which the hydrogen ions transmit, is extremely small, the generated gas becomes extremely minute bubble. And it is inferred that hydrogens in the state of minute bubble stick to portions contacting the membrane with the electrodes and a state is built up as if hydrogen membrane was formed. This phenomena resembles a phenomena called as polarization effect in the field of battery. As it is difficult for the minute bubbles of hydrogens to exfoliate by such physical impact as a stream, the resistance increases gradually and the membraneous ability deteriorates. And it is inferred that, the



Next, the second method of the present invention will be explained. The first method is a method wherein the recovery of the membrane function is carried out by changing the pressing force from the electrodes to the electrolyte membrane. On the other hand, in the second method, instead of the changing the pressing force, an operation for changing a current or voltage value is carried out.

Fig. 11 shows a running time chart concerning this second method. When the apparatus runs to control the electric current so that the concentration  $X$  of ozone aqua may be maintained at the predetermined desired value  $X_s$ , the current value  $A$  rises gradually according to the mentioning above. Therefore, the operation is started at the suitable time (time  $t_{51}$ ) before the current value  $A$  reaches the permission ceiling value  $A_e$  of the apparatus and carried out during a predetermined period ( $T_5$ ) from the time  $t_{51}$  to the time  $t_{52}$ , which changes the current value  $A$  between a permission maximum value  $A_e$  and a low current value  $A_2$  with pulses-shaped pattern. In this case, the voltage  $V$  varies between a low voltage value  $V_2$  and a permission maximum voltage value  $V_e$  of an apparatus with pulses-shaped pattern in reversibly to the change of the electric current  $A$ . On the other hand,

as to the concentration X of ozone aqua, only the slight change occurs and the big change does not occur. This pulse-shaped electric current change is repeated with plural times in a predetermined period (T5), the maximum value Ae and the minimum value A2 are maintained during a minute time ( $\Delta t$ ) respectively.

In this method, it is inferred that the membrane function of the electrolyte membrane 5 recovers to shatter the various ions or ion clusters accumulated in the membrane or the hydrogen minute bubbles generated between the membrane and the electrodes by means of electric shocks. The various ions, ion clusters or the hydrogen minute bubbles are caused from pressing the electrolyte membrane 5 continuously with the predetermined pressing force P1 as mentioned above.

In addition, as operations for varying the electric current, there are some following methods. The method for changing the electric current with pulse-shape pattern for the predetermined time T5 such as shown in Fig. 11. A method in which the current value changes into a low current value then the state is held for a regular time, after that, the current value changes into a high current value then the state is held for another regular time. In



the change of the electric current A or the voltage V reaches the limit, the various kinds of methods can be adopted such as follows.

1) The method of exchanging the electrolyte membrane to stop the apparatus.

2) The method explained in Fig. 8, in other words, the method in which, after the membrane is recovered to stop the apparatus and release the pressing force P of the electrodes 3,4 against the electrolyte membrane, the running is started again.

3) The method explained in Fig. 10, in other words, the method in which the revitalization of the membrane function is performed to raise the pressing force P against the electrolyte membrane step-by-step as function of electrolyte membrane deteriorates.

As mentioned above, in the first method, it is required that the employed ozone aqua production apparatus has a structure in which the anodic electrode 3, the cathode electrode 4 or the both electrodes can advance or retreat against the electrolyte membrane 5. There is the apparatus of Fig. 21 as an example of the ozone aqua production apparatus employed in the first method, However, the apparatus employed in the present invention is not limited to this. Therefore, the other optional ozone

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aqua production apparatus can be employed, which comprises the mechanism for advancing or reversing the electrodes . As concerns the ozone aqua production apparatus which comprises the mechanism for advancing or reversing the electrodes, other embodiment will be explained as follows.

Fig. 12 shows a feature section of the example ozone aqua production apparatus which comprises the mechanism for advancing or reversing the electrodes.

The electrolyte membrane 5 are arranged between the anode side casing 1 and the cathode side casing 2. The anodic electrode has a corrosion resistance as against ozone. Thereby, the anode chamber 6 and the cathode chamber 7 are respectively formed in anode side casing 1 and the cathode side casing 2 . The anodic electrode 3, which is provided with noble metals 16 having a catalyst function for generating ozone, is pressed onto one face of the electrolyte membrane 5 at the anode chamber side so as to be in contact each other. Additionally, the cathode electrode 4, which is provided with a contact face having noble metals 20, is pressed onto another face of the electrolyte membrane 5 at the cathode chamber 7 side so as to be in contact each other. The inflow openings 8,9 and outflow openings 10,11 of material water are formed by each of the anode chamber 6 and



the cathode chamber 7. DC voltage is applied between the both electrodes 3,4 by means of a direct-current power source 24 and electricity members 35,36. Additionally, an elastic rubber membrane 31 is arranged between a back face of an electrode plate 18 of the anodic electrode 3 and anode side casing 1. Thereby, the anode chamber 6 is made sealed. On the other hand, an elastic rubber membrane 32 is arranged between a back face of the cathode electrode 4 and the cathode side casing 2. Thus, the cathode chamber 7 is also made sealed. Air pressure or hydraulic pressure can be supplied from a pressure resource 41 to the anode chamber 6 and the cathode chamber 7 by means of pipes 37,38 connected with through-holes 33,34.

The anodic electrode 3 and the cathode electrode 4 advance toward the electrolyte membrane 5 by air pressure or hydraulic pressure respectively when the air pressure or the hydraulic pressure is supplied from the pressure resource 41 to the anode chamber 6 through the switch valve 39 and the pipe 37, and to the cathode chamber 7 through the a pipe 38. And then, the electrolyte membrane 5 is pressed at both sides. Operations for lowering the pressing forces of the both electrodes 3,4 against the electrolyte membrane 5 are carried out as follows.

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The pipes 37,38 are connected with the vent pipe 40 to control the switch valve 39. The hydraulic pressure or the air pressure in the anode chamber 6 and the cathode chamber 7 is ejected from the vent pipe 40. Then, the pressures in both chambers 6,7 fall down and also the pressing force falls down. In addition, the same parts as shown in the apparatus of Fig. 21 can be referred to the same symbol as shown in the apparatus of Fig. 21. So the detailed description has been omitted.

Next, Fig. 13 shows a feature section of other example apparatus having the mechanism for advancing or reversing the electrodes in the present invention. The apparatus of Fig. 12 is one of the example apparatuses with an air compressor. On the other hand, the ozone aqua production apparatus of Fig. 13 is one of the example apparatuses with a hydraulic machine which takes an advantage of material water. The ozone aqua production apparatus of Fig. 13 is basically same as the apparatus of Fig. 12 except that a hydraulic machine 51 is employed instead of the pressure resource 41 of Fig. 12. The same constitution as Fig. 12 can be referred to the same symbol in Fig. 12. So, an overlapping description has been omitted.

In Fig. 13, a pipe 55 which is branched off a

material water pipe, is connected to the hydraulic machine 51 at the positive pressure side of a cylinder chamber 53 of a large diameter piston 52. The pipe has a solenoid valve 54 on the way. On the other hand, a pipe 57 which is branched off a material water pipe, is also connected to the hydraulic machine 51 at the back pressure side. The pipe 57 also has a solenoid valve 56 on the way. Additionally, a drain pipe 61 is arranged between the solenoid valves 54,56 and the hydraulic machine 51, which lead to a drainage 60 through another solenoid valves 58,59. Further, at the positive pressure side of the cylinder chamber 63 of a small diameter piston 62, the pipes 65a, 65b are respectively connected to through-holes 33,34 of the anode chamber 6 and the cathode chamber 7 through non-return valve 64. In a bypassing circuit 67 on the non-return valve 64, a solenoid valve 66 is arranged.

Next, the run of this apparatus will be explained. At first, material water is supplied in a state in which the solenoid valves 54,59 are opened while the solenoid valves 56,58,66 are closed. Then, material water is supplied into the positive pressure side of the large diameter piston 52 of the hydraulic machine 51 through the solenoid valve 54

and the pipe 55. And, at the back pressure side of the large diameter cylinder chamber 53, the water is drained through the solenoid valve 59 and the pipe 61. Then, the large diameter piston 52 advances and increases the pressure of the pressure medium (water is preferable ) in the small diameter cylinder chamber 63. The pressure medium is supplied from the pipes 65a, 65b to the anode chamber 6 and the cathode chamber 7 through the non-return valve 64. And then, both electrodes advance by the back pressure and presses the electrolyte membrane 5 on the both sides. This pressing force is adjusted to change an opening degree of the solenoid valve 54. In this state, the material water flows in the apparatus through the inflow openings 8,9, then flows out at the outflow openings 10,11 through the passage consisted of lath nets 17,21. Accordingly, when both electrodes 3,4 are applied an electric current from the direct-current power source 24, ozone aqua generates at the anode side by the water electrolysis, then the ozone aqua flows out from the outflow opening 10. When predetermined time elapses, the solenoid valve 54 is closed and the solenoid valve 58 is opened. Then, the hydraulic pressure against the large diameter cylinder chamber 53 by material water is released. The large diameter

chamber cylinder 53 is opened toward the drainage 60 by through the pipes 55 and the solenoid valve 58. As a result, the pressure of the small diameter cylinder chamber 63 is also released. And the pressures acting on the anode chamber 6 and the cathode chamber 7 fall down. Then, the pressing force of the anodic electrode 3 and the cathode electrode 4 against the electrolyte membrane 5 falls down. With keeping this condition, after a regular time elapses, the solenoid valve 54 is opened and the solenoid valve 58 is closed as mentioned above, then the electrolyte membrane 5 is pressed by both electrodes 3,4 again. By means of repeating this operation of this solenoid valve requirement times, the above-mentioned operation for changing the pressing force is carried out.

In addition, in the present invention, as the advance or retreat mechanism of the both electrodes, any kind of types or any kind of structures can be employed. The advance or retreat mechanism of the both electrodes is not limited to the embodiment as mentioned above. Additionally, the shape of the electrodes and the structure are not limited to the embodiment illustrated in drawings. For example, the production efficiency of ozone aqua improves as follows: As illustrated in Fig. 14, as concerns width

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of the electrode plate, the electrodes are made up so that the anodic electrode 3 may be smaller than the cathode electrode 4. Accordingly, as concerns width of the chamber, the anode chamber 6 is made up so as to be narrower than the cathode chamber 7. In accordance with above structure, the current density at the anodic electrode 3 side becomes higher than one at the cathode electrode 4 side and the production efficiency of ozone aqua can be raised.

Additionally, the following other structure can be employed. As shown in Fig. 24, the structure in which only anodic electrode 3 is able to advance or retreat. As shown in Fig. 25, the structure in which only cathode electrode 4 is able to advance or retreat. In addition, in the above mentioned apparatus, the same constitution as in Fig. 14, Fig. 24 or Fig. 25 can be referred to the same symbol in those drawings. So, the detailed description has been omitted.

As discussed above, any shape and any structure of the apparatus can be employed in the present invention when the following condition is fulfilled. The following conditions is that at least one of the electrodes can advance and retreat against the electrolyte membrane and the pressing force of the electrodes against the electrolyte membrane is

adjustable by the electrodes. Accordingly, the present invention is not limited to the apparatuses and the methods as shown in the above embodiments.

Additionally, it is clear that the recovery method of the electrolyte membrane of the present invention can be applied as the recovery method about the electrolyte membrane of fuel cell. Because, the constitution of the electrolysis type ozone aqua production apparatus employed in the present invention is basically the same constitution as fuel cell.

(Example)

A concrete example of the present invention will be explained as follows. The examination which produces ozone aqua continuously, has been carried out with using the electrolysis type ozone aqua production apparatus having a constitution in which the anodic and cathode electrodes can be pressed on the electrolyte membrane by means of air pressure or hydraulic pressure as shown in Fig. 12. In addition, the electrolyte membrane used for the examination is a perfluorinated proton exchange membrane. Both anodic and cathode electrodes are made of the plate electrodes of 150cm<sup>2</sup>. Platina wire nets are respectively loaded on contact faces of the plate electrodes with the electrolyte

membrane, which have catalyst function for making generate ozone. On the back faces of the plate electrodes, lath nets made of titanium are arranged. Material water circulates in the lath net portions. The apparatus was run according to the method shown in Fig. 1 on the following operating condition.

[ Operating condition ]

- ① Set target ozone aqua concentration: 10ppm (mg/Liter)
- ② Initial value of the pressing force against the electrolyte membrane:  $P1=4\text{kg f/cm}^2$
- ③ Low value of the pressing force against the electrolyte membrane:  $P4=0\text{kgf/cm}^2$
- ④ Cycle of the operation for changing the pressing force:  $Tc = 15$  minute
- ⑤ The time which holds the pressing force at the low value pressure:  $T1=2$  second
- ⑥ Flow rate of ozone aqua: About 8 Liter / minute

In this running process, the current value A, the concentration X of ozone aqua, the voltage value V and the flow rate Q of ozone aqua were measured. The result was shown in Fig. 15. The pressing force was changed from  $4\text{kgf/cm}^2$  to 0 with 15 minutes cycle and the condition was held for two seconds. However, as is apparent from the figure above, the concentration of ozone aqua hardly changed and it was





for one month through the night and day without the stop caused from the membrane fouling as shown in the above-mentioned Fig. 8. And, as concerns the stability, the change was not recognized at all. Though the running was stopped after the one month passed, it was clear from this result of the test that further continuous running was possible.

(Comparative example)

As shown in Fig. 22, the running was carried out by means of the conventional method with use of the same apparatus as in the above example and the same condition as the above example except the operation for changing the pressing force was performed. The result is shown in Fig. 16. As was apparent from the figure above, the current value started rising when about 3 hours elapsed after start-up. When about 5 hours elapsed, the current value A reached the vicinity of the ceiling value of 180A of the apparatus. At this time, the concentration of ozone aqua was still maintained at the predetermined value of 10ppm. So, further continuous run was possible. However, as it was judged that the current value A reached the ceiling value soon, the apparatus stopped temporarily and the anodic electrode and the cathode electrode were kept apart from the electrolyte membrane. With this condition, the

apparatus was let alone for around 1 hour and 10 minutes. And, the function of the electrolyte membrane recovered. After that, with the initial operating condition, the running was begun again. Then, when about 5 hours elapsed in similar, the current value reached to the vicinity of the ceiling value of the apparatus. Incidentally, in the figure above, as concerns the electric current, the voltage and the flow rate, all values suddenly decreased during about 12:20-13:10 (while the ozone aqua concentration was generally uniformity). The reason for this is as follows. According to mentioning above, the material water was taken out from the tap of water supply in the office building. A consumption of tap water increased at the lunch break. Relating to it, a quantity of water supply to this apparatus decreased. The electric current and the voltage decreased so as to make the concentration of ozone aqua uniformity. Accordingly, as mentioned above, if the control for making the flow rate uniformity was carried out, such a decrease phenomena would disappear.

As is apparent from a comparison the example of the above present invention with the comparative example by the conventional method, in the present invention, the running can be continued for one month



pressing force of the electrodes against the electrolyte membrane is appropriately carried out without production stop. Then, the electrolyte membrane is recovered from on the way of the deterioration process. Additionally, according to the second method, while producing ozone aqua, the current value or the voltage value changes rapidly to increase or decrease. Then, the electrolyte membrane is recovered from on the way of the deterioration process. Because of this, in both method, the long term continuous running of the electrolysis type ozone aqua production apparatus becomes possible, which has been considered to be impossibility conventionally and the heart's desire for a person skilled in the art.

Especially, as is clear from the example, the present invention enables the stable continuous running for one month or more (it has not been examined yet until when the continuous running is possible). Therefore, it is a great effect as indescribable that the present invention enables to run continuously without stoppage, which is the heart's desire of a person skilled in the art and not exaggerated.

Further, even if such an operation as for changing the pressing force against the above

electrolyte membrane or for changing the electric current or the voltage forcibly is carried out, the recovery of the electrolyte membrane will reach a limit. At that time, if following things are carried out, it will become possible to extend the life time of the electrolyte membrane by leaps and bounds. The following things are that the electrolyte membrane is recovered to stop the apparatus temporarily and keep apart the electrodes from the electrolyte membrane. After that, the apparatus is run again. It cannot imagine how life extends thereby.

Additionally, if a following method for continuing the production of ozone aqua is employed, it will be possible that the running time of the ozone aqua production is extended by leaps and bounds beyond the comparison with the conventional method.

The following method is that the pressing force value is changed into the high value set previously when the electrolyte membrane will be not recovered enough even if the operation for changing the pressing force, the electric current or the voltage forcibly is carried out.

As a result, it becomes possible that a unit price of ozone aqua is largely reduced, too. Because, as concerns a rate of the operation of the

electrolysis ozone aqua production apparatus, a wonderful rate of the operation near to 100% is possible. The conventional rate was around 50%. Further, it is possible that the life time of the expensive electrolyte membrane is improved by leaps and bounds. Accordingly, it is possible to reduce the running cost of the apparatus and the ozone aqua product cost remarkably.

Additionally, the facilities becomes easy to be treated for a user and more spread of the ozone aqua production apparatus becomes possible. It is possible for the apparatus to be run continuously in the long term. As a result, the apparatus maintenance cost such as the exchange of the electrolyte membrane is reduced, too, in addition to the reduction of ozone production cost. And, the user is set free from the complexity of maintenance.

From the above, the factory cost of ozone aqua is reduced, as well, the ozone aqua production apparatus becomes easy to be treat for, even if, a user is an amateur in facilities. As a result, a disinfection or a cleaning by ozone aqua or other application of ozone aqua spreads out by leaps and bounds. From an environmental sanitation face of the nation, a big improvement and an effect are expected. The social spreading effect is

immeasurable.

#### Industrial applicability

As is explained in detail above, in the electrolysis type ozone aqua production method and the apparatus related to the present invention, it is possible to run continuously for a long time. Accordingly, in various fields of various cleaning and disinfection by ozone aqua, the present invention is extremely useful. Additionally, the recovery method of solid polymer electrolyte membrane of the present invention is the technology which can be applied to an apparatus of every field using a solid polymer electrolyte membrane. Especially, in the continuous running of the above electrolysis type ozone aqua production apparatus, the present recovery method is extremely useful technology.

00926344-101704